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Solubility product = K sp Consider the equilibrium between undissolved silver chloride and a saturated silver chloride solution: AgCl(s) \u2192 Ag+(aq)+Cl(aq) Since this is a heterogeneous equilibrium, the equilibrium constant is K sp = [Ag +][Cl] o always write the K sp expression with the solid on the left

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At a certain temperature, the solubility of Fe (OH) 2 in water is 7.7 x 10 -6 mol/L (M). Its K sp can be calculated based on the equilibrium equation: Fe(OH)2 \u2192 Fe2+ +2OH\u2192 Fe (OH) 2 \u2192 Fe 2 + + 2 OH \u2192. Therefore, the solubility product expression is: Ksp = [Fe2+][OH\u2192]2 K sp = [Fe 2 +] [OH \u2192] 2.

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