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Structures,

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Bases Study

An acid is a
substance that
contains

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hydrogen and
ionizes to
produce hydrogen
ions in aqueous
solution; A base
is a substance
that contains a
hydroxide group
and dissociates
to produce a
hydroxide ion in
aqueous
solution.

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associated with
acids, bases and
salts. Each
of...

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STUDY. PLAY.

Acids. taste
sour, change the
color of an acid
base indicator,
can be strong or
weak

electrolytes in
aqueous solution
Ex: citrus.

bases. taste
bitter, feel
slippery, will
change the color

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of an acid base indicator, can be strong or weak electrolytes in aqueous solution. Ex: soap

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"Acids, Bases,

and Salts". Use

these activities

to study the

vocabulary and

major concepts

presented in

this chapter.

compounds

derived from the

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reaction of a strong base with a weak acid or a strong acid with a weak base or a solution of a weak base and one of its salts. Quia - Chapter 19 "Acids, Bases, and Salts"

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Acids & Bases

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covers

properties of
acids and bases,

the different
definitions,

[H⁺] and [OH⁻]

and pH and pOH,

K_a and K_b and

their use when

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dealing with

weak acids and
bases and

buffers. Chapter

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Answer The
Arrhenius
definition states
that acids are
compounds that
contain
hydrogen, and
when placed in
aqueous solution
, ionize to form

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hydrogen (H⁺ ions). Chapter
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Bases, and Salts
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SECTION 19.1

ACID-BASE

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section compares and contrasts acids and bases as defined by the theories of Arrhenius, Brønsted-Lowry, and Lewis. It also identifies conjugate acid-base pairs in acid-base reactions.

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and Salts □□What

are the

properties of

acids and bases?

acids taste

sour, will

change the color

of an acid-base

indicator, and

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can be strong or weak

electrolytes in
Samples

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conjugate acid:
particle formed
when a base

gains a hydrogen
ion: conjugate

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base: particle remaining when an acid donates a hydrogen ion:

diprotic acid: acids containing two ionizable

hydrogens: Lewis acid: substance that accepts a pair of

electrons in a reaction: Lewis base: substance

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that donates a pair of electrons in a reaction:

Arrhenius acid: substances that release hydrogen ions (H^+) in aqueous solution:

Arrhenius base

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the world.

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STUDY. PLAY.

Tastes sour.

Acid. Changes

the color of an

acid-base

indicator (acid,

base, or both)

Both. Can be

strong or weak

electrolytes in

aqueous

solution. Page

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Acids and Bases

Section 18.1

Figure 18.1

Rhododendrons

flourish in

rich, moist soil

that is

moderately

acidic, whereas

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sempervivum, commonly called hen and chicks, grows best in drier, slightly basic soil.

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