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Solution:  $p_{\text{abs}} = p_{\text{atm}} + p_{\text{gage}} = 101.3 \text{ kPa} + 20 \text{ in. H}_2\text{O} = 124.8 \text{ kPa}$   
 $1 \text{ in. H}_2\text{O} = 2.488 \text{ kPa}$  Ideal gas law:

$\rho = \frac{p}{RT} = \frac{124.8 \text{ kPa}}{1 \text{ kg/K} \cdot 2077 \text{ J}} = 293.2 \text{ K}$   
 $1000 \text{ Pa} = 1 \text{ kPa}$   $\rho = 1 \text{ kg/m}^3$   $\rho = 0.181 \text{ kg/m}^3$  b.) Situation: A

sphere of 93 mm diameter contains an ideal gas.  $T = 20 \text{ C} = 293.2 \text{ K}$  Find: Calculate the density of argon at a vacuum pressure of 8.8 psi.

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